

# An elementary study of chemistry; introductory college course

## The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

**Groups (Columns)**  
The elements in each group have the same number of valence electrons.

Period	Group 1 (1 valence e <sup>-</sup> )	Group 2 (2 valence electrons)	Group 3 (3 valence electrons)	Group 4 (4 valence electrons)	Group 5 (5 valence electrons)	Group 6 (6 valence electrons)	Group 7 (7 valence electrons)	Group 8 (8 valence electrons)	Group 9 (9 valence electrons)	Group 10 (10 valence electrons)	Group 11 (11 valence electrons)	Group 12 (12 valence electrons)
1 <sup>st</sup> Period (1 shell)	1 H Hydrogen											2 He Helium
2 <sup>nd</sup> Period (2 shells)	3 Li Lithium	4 Be Beryllium	5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon				
3 <sup>rd</sup> Period (3 shells)	11 Na Sodium	12 Mg Magnesium	13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon				
4 <sup>th</sup> Period (4 shells)	19 K Potassium	20 Ca Calcium	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton				

**Periods (rows)**  
The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

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